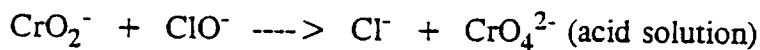


Electrochemistry #4

Key.

Use the following unbalanced equation to answer questions 1 - 9.



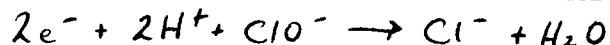
1. What is the highest oxidation number of chromium in the equation?

+6

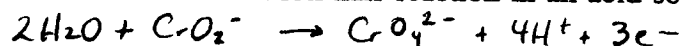
2. What is the highest oxidation number of chlorine in the equation?

+1

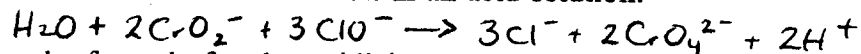
3. Write the balanced reduction half reaction in an acid solution.



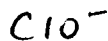
4. Write the balanced oxidation half reaction in an acid solution.



5. Write the balanced full reaction in an acid solution.



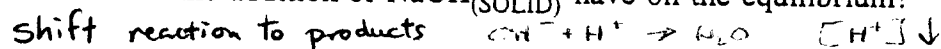
6. Write the formula for the oxidizing agent.



7. Write the formula of the species that is getting reduced.



8. What effect would the addition of NaOH(SOLID) have on the equilibrium?



9. What effect would the addition of NaOH(SOLID) have on the cell potential?

increases cell potential

eq^m shifts
→